

TITLE OF THE INVENTION

K-2060DIV

LOW TRANSMITTANCE GLASS

CROSS REFERENCE TO RELATED APPLICATION

5 This is a divisional application of a patent application
Serial No. 10/166,064 filed on June 11, 2002.

FIELD OF THE INVENTION

[0001] The present invention relates to a low transmittance
10 glass which has advantages of high heat ray absorptivity, high
quality, and high productivity. It particularly relates to a
glass having low visible light transmittance which has a smoky
color, such as gray to green, and has a high heat ray absorption
performance, it more particularly relates to a low transmittance
15 glass which is suitable for a window glass to be reinforced and
installed in a vehicle.

[0002] The present invention relates to a green glass which
has advantages of high heat ray absorptivity, high ultraviolet
absorptivity, high quality, and high productivity. It
20 particularly relates to a glass having high visible light
transmittance which has a shade of green, and has a high heat
ray absorption performance, and it more particularly relates to
a green glass which is suitable for a window glass to be
reinforced and installed in a vehicle.

25

BACKGROUND OF THE INVENTION

[0003] A variety of glasses with ultraviolet/infrared
absorptivity to be used as a vehicle windshield have been
proposed with the view of preventing degradation of luxurious

interior materials and reducing cooling load of the vehicle. In view of comfort of passengers and privacy protection, a glass having low visible light transmittance is preferably used for a rear window glass of a vehicle.

5 [0004] The front windshield of a vehicle is obligated to have a visible light transmittance higher than a specific level for enough visibility of a driver. A glass having high visible light transmittance and provided with ultraviolet and heat ray absorptivity has a greenish shade because the ends of its
10 ultraviolet absorption range and infrared absorption range overlap the visible range.

[0005] Ultraviolet/infrared absorbent glasses having low to middle visible light transmittance, low ultraviolet transmittance and low total solar radiation transmittance have
15 been disclosed in Japanese patent H10-114540A and H10-45425A. Since these ultraviolet/infrared absorbent glasses have high heat ray absorption characteristics, productivity of each glass in a glass melting furnace is low. Inside the melting furnace, a top of the glass is directly heated with flames, but the
20 bottom of the glass material can not be sufficiently heated because a large part of heat rays directed at the surface of the glass material by the radiation of the flames are absorbed by the top of the glass material. It is thus difficult to melt the glass material in the melting furnace uniformly. . "

25 [0006] The glass material is necessarily maintained at a higher-than-normal degree of reduction in order to provide high heat ray absorption characteristics thereto. Therefore, a large amount of reducing agents, mainly including graphite powder, etc., can be added into the glass batch, but such agents are

liable to cause unfused silica because the agents can exceedingly react with sulfate, mainly including sodium sulfate, etc., which is added to the glass batch as a refining agent.

[0007] In order to melt the glass material uniformly, at least one of the following processes can be adopted, such as: lowering the amount of the glass material below that for the ordinary operation; heating the bottom of the glass material by energizing electrodes inserted into the bottom of the furnace; and bubbling the glass material.

[0008] However, lowering the amount of the glass material below that for the ordinary operation causes decrease of the production capacity, thereby raising the production cost. Electrical heating by insertion of electrodes in the bottom of the furnace and bubbling the glass material require modification of the facilities. These processes can cause defects, such as: generation of a lot of defects such as bubbles in the glass depending upon the operating conditions, and result in significant decline of the production capacity.

20 SUMMARY OF THE INVENTION

[0009] The low transmittance glass of the present invention is composed of soda-lime-silica glass comprising 0.001 to 2 wt.% Li₂O and, as colorant, 0.7 to 2.2 wt.% total iron oxide (T-Fe₂O₃) expressed as Fe₂O₃. The glass with a thickness between 2.1 mm and 6mm has visible light transmittance (YA) of not greater than 65%, measured by using the CIE illuminant A, total solar radiation transmittance (TG) of not greater than 60%, and ultraviolet transmittance (Tuv) defined by ISO 9050 of not greater than 25%.

[0010] The low transmittance glass of the present invention has superior heat ray absorption properties, and it is improved in quality and productivity by including Li₂O in its base glass composition so as to lower viscosity of the glass material, 5 accelerating melt and homogenization of the glass material, and it also has superior capacity for reinforcement.

[0011] Since the glass material of the low transmittance glass of the present invention has low viscosity, melt and homogenization of the glass material are accelerated. The low 10 transmittance glass is improved in quality and productivity, and is provided with low infrared transmittance.

[0012] Since when applied with reinforcement by air blast cooling, the low transmittance glass of the present invention obtains higher surface compression than that of conventional 15 ones, it is superior in capacity for reinforcement. The low transmittance glass of the present invention has low visible light transmittance and low ultraviolet transmittance, so that it is suitable for a rear view window of a vehicle.

[0013] The ultraviolet/infrared absorbent green glass of the 20 present invention is composed of soda-lime-silica glass comprising 0.001 to 2 wt.% Li₂O and, as colorant, 0.4 to 2 wt.% total iron oxide (T-Fe₂O₃) expressed as Fe₂O₃ wherein FeO expressed as Fe₂O₃ is 15 to 60% of T-Fe₂O₃. The glass with a thickness between 2.1 mm and 6 mm has visible light 25 transmittance (YA) of not less than 70%, measured by using the CIE illuminant A, and total solar radiation transmittance (TG) of not greater than 60%.

[0014] The ultraviolet/infrared absorbent green glass of the present invention has superior heat ray absorption properties,

and it is improved in quality and productivity by including Li₂O in its base glass composition so as to lower viscosity of the glass material, accelerating melt and homogenization of the glass material, and it also has superior capacity for reinforcement.

[0015] Since the glass material of the low ultraviolet/infrared absorbent green glass of the present invention has low viscosity, melt and homogenization of the glass material are accelerated. The green glass is improved in quality and productivity, and it is provided with low infrared transmittance. Since when applied with reinforcement by air blast cooling, the glass obtains higher surface compression than that of conventional ones, it is superior in capacity for reinforcement. The green glass has high visible light transmittance and low ultraviolet transmittance, so that it is suited for a window of a vehicle.

BRIEF DESCRIPTION OF DRAWINGS

[0016] Fig.1 is a graph showing a relationship between Li₂O content and temperature at which log η becomes 2;

Fig 2 is a graph showing a relationship between Li₂O content and surface compression; and

Fig 3 is a graph showing a relationship between Li₂O content and transmittance of light having a wavelength of 700 nm.

25

DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENTS

[Low transmittance glass]

[0017] A low transmittance glass of the present invention may be composed of a base glass composition comprising:

65 to 80 wt.% SiO₂;

0 to 5 wt.% Al₂O₃;

0 to 10 wt.% MgO;

5 to 15 wt.% CaO wherein a total amount of MgO and CaO is

5 5 to 15 wt.%; 10 to 20 wt.% Na₂O;

0 to 5 wt.% K₂O wherein a total amount of Na₂O and K₂O is
10 to 20 wt.%; and

0 to 5 wt.% B₂O₃.

[0018] The low transmittance glass of the present invention
10 is preferable to include at least one selected from the group
consisting of Se, CoO, Cr₂O₃, and NiO as colorant. The
preferable content of Se is 0.0001 to 0.1 wt.%, that of CoO is
0.0001 to 0.1 wt.%, that of Cr₂O₃ is 0.001 to 2 wt.%, and that of
NiO is 0.001 to 2 wt.%.

15 [0019] The low transmittance glass of the present invention
is preferable to include at least one selected from the group
consisting of TiO₂, CeO₂, MoO₃, V₂O₅ and La₂O₃ as colorant in an
amount of 0.0001 to 1 wt %.

[0020] This low transmittance glass of the present invention
20 has visible light transmittance (YA) of 5 to 65%, measured by
using CIE illuminant A, total solar radiation transmittance (TG)
of not greater than 50%, and ultraviolet transmittance (Tuv)
defined by ISO 9050 not greater than 20%, in case that the glass
has a thickness between 2.1 mm and 6 mm.

25 [0021] The description will be made as regard to the low
transmittance glass composition of the present invention. It
should be noted that content of each component will be
represented with percentage by weight.

[0022] Li₂O is a component for lowering viscosity of the glass material and thus improving productivity of the glass. Since Li₂O has an effect of making thermal expansion coefficient and Young's modulus of the glass larger in case that the content of 5 Li₂O is in a specific range, Li₂O is also a component for increasing surface compression of the glass when the glass is applied with reinforcement by air blast cooling, improving a capacity for reinforcement of the glass.

[0023] When the content of Li₂O is not greater than 2%, 10 viscosity of the glass material becomes lowered as the content of Li₂O increases, and thus productivity of the glass becomes improved. However, when the content of Li₂O is greater than 2%, the effect of lowering viscosity decreases even though the content of Li₂O increases. When the content of Li₂O is less than 15 0.001%, Li₂O does not sufficiently give its effect of lowering viscosity to the glass material. The relationship between the content of Li₂O and the temperature at which log η becomes 2 in a typical soda-lime-silica glass is shown in Fig.1. It is noted that the temperature at which log η becomes 2 decreases 20 monotonously as the content of Li₂O increases, and the slope becomes gentle as the content of Li₂O exceeds 2%. Therefore, the content of Li₂O is set in the 0.001 to 2% range in the present invention. Since Li₂O material is costly, the content of Li₂O is preferably taken as less than 2% from a compromise between the 25 cost and the effect of Li₂O.

[0024] When the content of Li₂O is not greater than 0.5%, surface compression of the glass reinforced with air blast cooling increases as the content of Li₂O increases. When the content of Li₂O is more than 0.5%, surface compression of the

glass is approximately constant even though the content of Li₂O increases. The relationship between the content of Li₂O and the surface compression in a typical soda-lime-silica glass is shown in Fig. 2. It is noted that although the surface compression 5 increases as the content of Li₂O increases, it becomes approximately constant as the content of Li₂O exceeds 0.5%. The content of Li₂O is preferably taken as 0.5% or less from a compromise between the cost and the effect of Li₂O. When the increase of surface compression of the glass by addition of Li₂O 10 is expected, Li₂O is preferably added in an amount of 0.05% or more.

[0025] Iron oxide is present in the form of Fe₂O₃ and the form of FeO in the glass. Fe₂O₃ is a component for improving the ultraviolet absorptivity and FeO is a component for improving 15 the infrared absorptivity. When the total amount of iron oxide (T-Fe₂O₃) expressed as Fe₂O₃ is less than 0.7%, the efficiency of ultraviolet and infrared absorptivity becomes small so as not to provide desired optical properties. On the other hand, when T-Fe₂O₃ exceeds 2.2%, the infrared absorptivity of T-Fe₂O₃ becomes 20 too high to produce the glass in an ordinary melting furnace, and the color unpreferably becomes too greenish. In case of successively producing glasses by a glass melting furnace with a large amount of T-Fe₂O₃, long time is required to change the glass composition in the furnace.

[0026] In the present invention, an effect of shifting the light absorption peak of FeO toward the short wavelength range of light is given by coexistence of FeO and Li₂O in the soda lime glass. In order to illustrate this effect, the relationship 25 between the content of Li₂O and the transmittance of light having

a wavelength of 700 nm is shown in Fig. 3. It is noted that the transmittance of light having a wavelength of 700nm decreases according as the content of Li_2O increases. Addition of Li_2O makes it possible to lower visible light transmittance effectively with a small content of FeO .

[0027] When the $\text{FeO}/\text{T-Fe}_2\text{O}_3$ ratio (a weight of FeO expressed as Fe_2O_3 against $\text{T-Fe}_2\text{O}_3$) is less than 20%, sufficient heat ray absorptivity can not be obtained. When $\text{FeO}/\text{T-Fe}_2\text{O}_3$ ratio is more than 50%, silica-rich ream and silica scum are present in a glass because the glass is highly reduced, resulting in decrease of productivity.

[0028] SiO_2 is a main component for forming skeleton of glass. Less than 65% SiO_2 lowers the durability of the glass and more than 80% SiO_2 raises the melting temperature of the glass so high.

[0029] Al_2O_3 is a component for improving the durability of the glass. More than 5% Al_2O_3 raises the melting temperature of the glass so high.

[0030] MgO and CaO improve the durability of the glass and adjust a liquidus temperature and viscosity of the glass. More than 10% MgO raises the liquidus temperature. Less than 5% or more than 15% CaO raises the liquidus temperature of the glass. The durability of the glass is lowered when the total amount of MgO and CaO is less than 5%, while the liquidus temperature is increased when the total exceeds 15%.

[0031] Na_2O and K_2O prompt the glass to melt. The efficiency of promotion of melting becomes poor when Na_2O is less than 10% or the total of Na_2O and K_2O is less than 10%, while the durability of the glass is lowered when Na_2O exceeds 20% or the

total of Na₂O and K₂O exceeds 20%. K₂O is preferable not to exceed 5% because of its expensive cost.

[0032] B₂O₃ is a component for improving the durability of the glass, prompting to melt, and yet enhancing the ultraviolet absorption. B₂O₃ should be less than 5%, since difficulties during molding are caused due to the vaporization of B₂O₃ when B₂O₃ exceeds 5%.

[0033] Se, CoO, Cr₂O₃ and NiO are components for adjusting visible light transmittance and color of the glass, so that it is preferable to add at least one of these components in the glass.

[0034] Se gives a red to pink shade to the glass and also gives grayish shade to the glass by cooperating with FeO or CoO. When Se is included in the glass, more than 0.1% Se reduces visible light transmittance too much, so that the glass can not be provided with desired properties. The preferable content of Se is 0.0001 to 0.1%.

[0035] CoO gives a blue shade to the glass and also gives grayish shade to the glass by cooperating with Se, NiO or Fe₂O₃. When CoO is included in the glass, more than 0.1% CoO reduces visible light transmittance too much, so that the glass can not be provided with desired properties. The preferable content of CoO is 0.0001 to 0.1%.

[0036] Cr₂O₃ gives a green shade to the glass and also adjusts visible light transmittance and color of the glass by cooperating with Se, NiO, CoO or Fe₂O₃. More than 2% CrO₃ reduces visible light transmittance too much, so that the glass can not be provided with desired properties. The preferable content of Cr₂O₃ is 0.001 to 2%.

[0037] NiO gives a brown to purple shade to the glass and also gives grayish shade to the glass by cooperating with FeO or CoO. More than 2% NiO reduces visible light transmittance too much, so that the glass can not be provided with desired optical properties, and it unpreferably activates formation of nickel sulfide stones. The preferable content of NiO is 0.001 to 2%.

[0038] In order to obtain more desirable shade and properties, at least one selected from the group consisting of TiO₂, CeO₂, MoO₃, V₂O₅ and La₂O₃ may be added as auxiliary ultraviolet 10 absorbing agent in an amount of 0.0001 to 1%.

[0039] Sulfate of alkaline or alkaline earth metal has been added as a refining agent for the glass, and the glass usually includes SO₃ in an amount of about 0.1 to 0.5%. One or more than two among Sb₂O₃, SnO₂, and the like may be added as a reducing 15 agent or a refining agent for the glass in an amount not greater than 1%. In order further securely to prevent the formation of nickel sulfide stones, ZnO may be added in an amount not greater than 1%.

[0040] The low transmittance glass of the present invention 20 has visible light transmittance (YA) of not greater than 65%, measured by using the CIE illuminant A, total solar radiation transmittance (TG) of not greater than 60%, and ultraviolet transmittance (Tuv) defined by ISO 9050 of not greater than 25%, when the glass has a thickness in a range of 2.1 to 6 mm. YA is 25 preferable to be in a range of 5 to 65%, TG is preferable not to exceed 50%, and Tuv is preferable not to exceed 20%.

[0041] Hereinafter, examples and comparative examples of the low transmittance glass of the present invention will be described.

[Examples 1-10]

[0042] The formulation of typical soda-lime-silica glass material is shown in Table 1. Lithium oxide, ferric oxide, metallic selenium, cobalt oxide, chromium oxide, nickel oxide, 5 titanium oxide, cerium oxide, molybdenum oxide, vanadium pentoxide and lanthanum oxide were added to the glass material as desired. The glass material thus prepared was held in an electric furnace at 1500°C for 4 hours. The molten glass was cast on a stainless plate and held at 650°C for 1 hour, and then 10 annealed to the room temperature in the furnace so as to obtain 6 mm thick glass plates.

Table 1

component	content [g]
silica sand	973.6
dolomite	255.7
limestone	30.24
soda ash	230.02
salt cake	9.881
carbon	0.617
total	1500

20

[0043] The obtained glass plates were polished, so that each glass plate has a thickness between 3.5 mm and 5 mm, so as to obtain the samples. Each sample was determined in optical 25 properties including visible light transmittance (YA) measured by using the CIE illuminant A, total solar radiation transmittance (TG), ultraviolet transmittance (Tuv) defined by ISO 9050, dominant wavelength (DW) and excitation purity (Pe) measured by using the CIE illuminant C.

[0044] The samples were also determined in physical properties including glass transition temperature (T_g), deformation temperature (T_d), mean thermal expansion coefficient of the 50 to 350°C range ($\alpha_{(50-350)}$), Young's modulus (E), surface compression and the temperature at which $\log \eta$ becomes 2. Each sample was formed into a rod-like shape having a length of 15 mm and a diameter of 5 mm, and then it was loaded with a load of 5 g and heated from the room temperature to 700°C at a rate of 10°C per minutes with silica glass rod prepared as a standard sample by the use of EXSTAR 6000 (SEIKO Electronics Inc.), so as to measure glass transition temperature (T_g), deformation temperature (T_d) and mean thermal expansion coefficient of the 50 to 350°C range ($\alpha_{(50-350)}$). Young's modulus (E) was measured with ultrasonic wave according to sing-around method, surface compression was measured with a Babinet-Style surface stress meter according to via-scope method, and the temperature at which $\log \eta$ becomes 2 was measured according to platinum ball drawing method.

[0045] The base glass composition, the content of colorant, $FeO/T-Fe_2O_3$ ratio, the optical properties and the physical properties of each sample were shown in Tables 2 and 3. The contents of the components in these tables are indicated as percentage by weight.

Table 2

		Example 1	Example 2	Example 3	Example 4	Example 5
glass composition [wt %]	SiO ₂	71.0	71.5	70.7	71.4	71.2
	Al ₂ O ₃	1.40	0.11	0.89	1.35	1.35
	MgO	3.64	3.75	4.13	3.68	3.68
	CaO	7.66	8.62	8.38	7.75	7.74
	Na ₂ O	13.7	13.6	13.2	13.7	13.6
	K ₂ O	0.91	0.07	0.23	0.77	0.76
	Li ₂ O	1.00	2.00	0.50	0.05	0.12
	T-Fe ₂ O ₃	1.25	1.40	1.42	1.30	1.30
	FeO/T-Fe ₂ O ₃	0.23	0.24	0.22	0.23	0.23
	NiO	0.065	—	—	0.016	0.098
optical properties	CoO	0.019	0.022	0.012	0.004	0.020
	Se	0.001	0.003	0.0005	—	—
	Cr ₂ O ₃	—	—	0.002	—	—
	TiO ₂	0.02	0.03	0.03	0.03	0.03
	thickness[mm]	5	4	4	4	4
	YA[%]	17.2	16.3	37.8	48.6	16.9
physical properties	TG[%]	17.0	13.7	27.5	27.0	15.4
	Tuv[%]	9.0	2.4	4.7	8.7	8.2
	DW[nm]	556	522	494	502	521
	Pe[%]	9.0	1.0	5.8	5.0	4.5
	Tg [°C]	531.6	499.1	531.9	547.1	544.8
	Td [°C]	603.2	557.7	602.7	625.4	623.7
	$\alpha_{(50-350)} \times e^{-7}$	93.4	92.6	95.8	90.0	94.4
	E[GPa]	76.2	77.3	74.3	74.0	75.0
	surface compressive stress[MPa]	112.0	112.8	112.0	109.0	111.3
	temperature at which log η becomes 2 [°C]	1397	1387	1419	1430	1431

Table 3

		Example 6	Example 7	Example 8	Example 9	Example 10
glass composition [wt.%]	SiO ₂	71.4	71.2	71.6	72.0	71.6
	Al ₂ O ₃	1.35	1.35	1.35	1.35	1.34
	MgO	3.69	3.68	3.70	3.72	3.70
	CaO	7.76	7.74	7.78	7.83	7.78
	Na ₂ O	13.3	13.3	12.4	11.4	11.3
	K ₂ O	0.77	0.76	0.77	0.77	0.77
	Li ₂ O	0.33	0.33	1.01	1.50	2.00
	T-Fe ₂ O ₃	1.30	1.30	1.31	1.31	1.31
	FeO/T-Fe ₂ O ₃	0.24	0.22	0.22	0.22	0.22
	NiO	0.016	0.098	0.098	0.099	0.099
optical properties	CoO	0.004	0.020	0.020	0.020	0.020
	Se	—	—	—	—	—
	Cr ₂ O ₃	—	—	—	—	—
	TiO ₂	0.03	0.03	0.03	0.03	0.03
	thickness[mm]	4	3.5	3.5	3.5	3.5
	YA[%]	49.8	20.8	20.7	20.6	20.6
physical properties	TG[%]	27.9	19.3	19.1	19.1	19.0
	Tuv[%]	9.8	10.3	10.1	9.8	9.8
	DW[nm]	502	525	524	523	524
	Pe[%]	5.0	4.3	4.3	4.4	4.2
	Tg [°C]	535.5	539.2	518.6	510.2	504.7
	Td [°C]	612.2	615.4	596.8	598.1	586.2
$\alpha_{(50-350)} \times e^{-7}$		95.1	96.5	91.2	90.2	93.8
E[GPa]		74.9	75.0	76.3	77.5	78.5
surface compressive stress[MPa]		112.1	114.3	109	110	117
temperature at which log η becomes 2 [°C]		1417	1429	1429	1396	1390

[0046] It is apparent from Tables 2, 3 that each glass of the present invention has the temperature at which $\log \eta$ becomes 2, which is an index of high temperature viscosity, so that the low transmittance glass of the present invention is excellent in 5 meltability. Each low transmittance glass of the invention has also high surface compression, so that the low transmittance glass has large capacity for reinforcement.

[0047] Examples 1-10 have preferable optical properties, such as: visible light transmittance (YA) measured by using the CIE 10 illuminant A of 5 to 65%, total solar radiation transmittance (TG) of not greater than 50% and ultraviolet transmittance (Tuv) defined by ISO 9050 of not greater than 20%. These examples also have dominant wavelength (DW) measured by using the CIE illuminant C of 490 to 560 nm and excitation purity (Pe) 15 measured by using the CIE illuminant C of not greater than 9%.

[Comparative Examples 1 and 2]

[0048] The glass composition, the optical properties and the physical properties of Comparative Examples 1 and 2 each of which was prepared in the same way as Examples 1-10 except for 20 the glass composition are shown in Table 4.

Table 4

		Comparative Example 1	Comparative Example 2
glass composition [wt.%]	SiO ₂	70.5	70.6
	Al ₂ O ₃	1.66	1.66
	MgO	3.72	3.70
	CaO	7.96	7.86
	Na ₂ O	13.9	13.7
	K ₂ O	0.79	0.81
	Li ₂ O	0.00	0.00
	T-Fe ₂ O ₃	1.26	1.29
	FeO/T-Fe ₂ O ₃	0.27	0.23
	NiO	0.016	0.064
	CoO	0.005	0.018
	Se	—	—
	Cr ₂ O ₃	—	—
optical properties	TiO ₂	0.03	0.03
	thickness[mm]	5.0	4.0
	YA[%]	39.9	23.3
	TG[%]	21.7	18.7
	Tuv[%]	5.8	8.1
	DW[nm]	502	496
physical properties	Pe[%]	5.9	7.1
	Tg [°C]	558.5	557.3
	Td [°C]	643.5	642.3
	$\alpha_{(50-350)} \times c^{-7}$	92.6	92.6
	E[GPa]	74.2	74.2
	surface compressive stress[MPa]	106.0	107.3
temperature at which log η becomes 2 [°C]		1436	1434

[0049] Comparative Examples 1 and 2 are conventional low transmittance glasses which do not include Li₂O. These glasses have higher temperature at which log η becomes 2 and lower surface compression than those of the present invention, 5 resulting in reduction of productivity and capacity for reinforcement.

[0050] As described above, the present invention provides a low transmittance glass composition having superior heat ray absorption characteristics and an advantage of high productivity. 10 Particularly, the present invention provides a glass having low visible light transmittance which has smoky color, such as gray to green, and high heat ray absorption characteristics, more particularly, it provides a low transmittance glass which is suited for a window glass to be reinforced and installed in a 15 vehicle.

[Ultraviolet/infrared absorbent green glass]

[0051] The ultraviolet/infrared absorbent green glass of the present invention preferably has the same base glass composition 20 as the low transmittance glass mentioned above.

[0052] The ultraviolet/infrared absorbent green glass of the present invention preferably includes as colorant:

0.4 to 1 wt.% total iron oxide expressed as Fe₂O₃ (T-Fe₂O₃) ;
0 to 1 wt.% TiO₂; and
25 0 to 2 wt.% CeO₂.

[0053] The ultraviolet/infrared absorbent green glass preferably has ultraviolet transmittance (Tuv) defined by ISO 9050 of not greater than 35%, dominant wavelength (DW) measured by using the CIE illuminant C of 490 to 560 nm, and excitation

purity (P_e) measured by using the CIE illuminant C of less than 6% when the glass has a thickness between 2.1 mm and 6 mm.

[0054] The ultraviolet/infrared absorbent green glass is preferable to include at least one selected from the group 5 consisting of Se, CoO, Cr₂O₃, Mn₂O₃, CuO, Nd₂O₃, Er₂O₃, MoO₃, V₂O₅ and La₂O₃ in an amount of 0.0001 to 0.1wt.%.

[0055] The description will be made as regard to the ultraviolet/infrared absorbent green glass composition of the present invention. It should be noted that content of each 10 component will be represented with percentage by weight.

[0056] The description of the reason why the above base glass composition is desirable for the low transmittance glass applies equally to that for the ultraviolet/infrared absorbent green glass.

15 [0057] The description will be made as regard to the colorant of the ultraviolet/infrared absorbent green glass of the present invention.

[0058] Iron oxide is present in the form of Fe₂O₃ and the form of FeO in the glass. Fe₂O₃ is a component for improving the 20 ultraviolet absorptivity and FeO is a component for improving the infrared absorptivity. When the total amount of iron oxide (T-Fe₂O₃) expressed as Fe₂O₃ is less than 0.4%, the efficiency of ultraviolet and infrared absorptivity becomes small so as not to provide desired optical properties. On the other hand, when T- 25 Fe₂O₃ exceeds 2%, visible light transmittance becomes too low and the color of the glass unpreferably becomes too greenish. When TiO₂ and CeO₂ also are included in the glass, the content of T-Fe₂O₃ is preferable not to be greater than 1%.

[0059] When the FeO/T- Fe_2O_3 ratio (a weight of FeO expressed as Fe_2O_3 against T- Fe_2O_3) is less than 15%, sufficient heat ray absorptivity can not be obtained. If FeO/T- Fe_2O_3 ratio is more than 60%, silica-rich ream and silica scum are present in a glass because the glass is highly reduced, unpreferably resulting in decrease of productivity and lapse of the effect of Li_2O . The FeO/T- Fe_2O_3 ratio is preferable not to be greater than 50%, more preferable not to be greater than 35%.

[0060] TiO_2 is a component for improving the ultraviolet absorptivity particularly by interaction with FeO. The content of TiO_2 is preferable not to be greater than 1%. More than 1% TiO_2 unpreferably raises cost and makes the shade of the glass yellowish. When the effect of absorbing ultraviolet of TiO_2 is expected, TiO_2 is preferably added in an amount of 0.01% or more.

[0061] CeO_2 is a component for improving the ultraviolet absorptivity and is present in the form of Ce^{3+} or in the form of Ce^{4+} in the glass. Particularly, Ce^{3+} is effective in absorbing ultraviolet with less absorptivity in the visible range. The content of CeO_2 is preferable not to be greater than 2%. More than 2% CeO_2 unpreferably raises cost and reduces visible light transmittance. When the effect of absorbing ultraviolet of CeO_2 is expected, CeO_2 is preferably added in an amount of 0.01% or more.

[0062] In order to obtain more desirable shade and optical properties, at least one selected from the group consisting of Se , CoO , Cr_2O_3 , Mn_2O_3 , CuO , Nd_2O_3 and Er_2O_3 as colorant and the group consisting of MoO_3 , V_2O_5 and La_2O_3 as auxiliary ultraviolet absorbing agent may be added in an amount of 0.0001 to 1%.

[0063] Sulfate of alkaline or alkaline earth metal has been added as a refining agent for the glass, and the glass usually includes SO₃ in an amount of about 0.1 to 0.5%. One or more than two among Sb₂O₃, SnO₂, and the like may be added as a reducing agent or a refining agent for the glass in an amount not greater than 1%. In order further securely to prevent the formation of nickel sulfide stones, ZnO may be added in an amount not greater than 1%.

[0064] The ultraviolet/infrared absorbent green glass of the present invention is preferable to have visible light transmittance (YA) of not less than 70%, measured by using the CIE illuminant A, and total solar radiation transmittance (TG) of not greater than 60% when the glass has a thickness between 2.1 mm and 6 mm.

[0065] The ultraviolet/infrared absorbent green glass of the present invention has ultraviolet transmittance (Tuv) defined by ISO 9050 of not greater than 35%, dominant wavelength (DW) measured by using the CIE illuminant C of 490 to 560 nm, and excitation purity (Pe) measured by using the CIE illuminant C of less than 6%.

[0066] Hereinafter, examples and comparative examples of the ultraviolet/infrared absorbent green glass of the present invention will be described.

[Examples 11-22]

[0067] Lithium oxide, ferric oxide, titanium oxide, cerium oxide, metallic selenium, cobalt oxide, nickel oxide, chromium oxide, manganese oxide, copper oxide, neodymium oxide, erbium oxide, molybdenum oxide, vanadium pentoxide, lanthanum oxide and carbon based reducing agent, including carbon powder, etc., were

added to the typical soda lime silica glass batch material as desired. The glass material thus prepared was held in an electric furnace at 1500°C for 4 hours. The molten glass was cast on a stainless plate, and then annealed to the room
5 temperature so as to obtain 6 mm thick glass plates.

[0068] The obtained glass plates were polished, so that each glass plate has a thickness between 2.6 mm and 5 mm, so as to obtain the samples. Each sample was determined in optical properties including visible light transmittance (YA) measured
10 by using the CIE illuminant A, total solar radiation transmittance (TG), ultraviolet transmittance (Tuv) defined by ISO 9050, dominant wavelength (DW) and excitation purity (Pe) measured by using the CIE illuminant C.

[0069] The samples were also determined in physical
15 properties including glass transition temperature (Tg), deformation temperature (Td), mean thermal expansion coefficient of the 50 to 350°C range ($\alpha_{(50-350)}$), Young's modulus (E), surface compression and the temperature at which $\log \eta$ becomes 2. Each sample was formed into a rod-like shape having a length of 15 mm
20 and a diameter of 5 mm, and then it was loaded with a load of 5 g and heated from the room temperature to 700°C at a rate of 10°C per minutes with silica glass rod prepared as a standard sample by the use of EXSTAR 6000 (SEIKO Electronics Inc.), so as to measure glass transition temperature (Tg), deformation
25 temperature (Td) and mean thermal expansion coefficient of the 50 to 350°C range ($\alpha_{(50-350)}$). Young's modulus (E) was measured with ultrasonic wave according to sing-around method, surface compression was measured with a Babinet-Style surface stress meter according to via-scope method, and the temperature at

which $\log \eta$ becomes 2 was measured according to platinum ball drawing method.

[0070] The glass composition and $\text{FeO}/\text{T-Fe}_2\text{O}_3$ ratio of each sample were shown in Tables 5 and 6. The contents of the components in these tables are indicated as percentage by weight. Tables 5 and 6 also show the optical properties and the physical properties of each sample.

Table 5

		Example 11	Example 12	Example 13	Example 14	Example 15	Example 16
glass composition [wt.%]	SiO ₂	70.8	70.9	69.5	69.5	68.2	68.3
	Al ₂ O ₃	1.74	1.88	1.71	1.91	2.87	2.87
	MgO	4.09	4.02	3.19	3.11	1.64	1.65
	CaO	7.94	7.96	8.26	8.44	9.21	9.22
	Na ₂ O	13.8	13.6	14.5	14.4	15.0	14.8
	K ₂ O	0.80	0.96	0.23	0.07	0.92	0.92
	Li ₂ O	0.05	0.09	0.06	0.12	0.16	0.33
	T-Fe ₂ O ₃	0.51	0.55	0.63	0.65	0.64	0.64
	FeO/T-Fe ₂ O ₃	0.23	0.24	0.32	0.31	0.22	0.18
	TiO ₂	—	—	0.09	0.14	0.16	0.16
optical properties	CeO ₂	—	—	1.65	1.65	1.17	1.17
	thickness[mm]	5.0	3.5	4.0	3.4	3.1	3.1
	YA[%]	78.5	80.3	72.9	75.3	77.0	77.5
	TG[%]	54.9	59.7	45.5	49.5	49.6	50.7
	Tuv[%]	24.5	32.4	8.6	10.1	15.0	15.3
	DW[nm]	500	500	508	510	497	498
physical properties	Pe[%]	2.4	1.9	2.6	2.2	3.1	2.9
	Tg [°C]	557.3	555.3	556.2	552.3	544.0	537.6
	Td [°C]	626.8	623.0	629.9	627.9	624.3	618.1
	$\alpha_{(50-350)} \times e^{-7}$	92.9	93.1	94.3	93.9	100.8	101.8
	E[GPa]	75.40	75.60	75.14	75.47	75.10	75.37
	surface compressive stress[MPa]	109.9	110.5	111.4	111.4	120.8	122.7
temperature at which log η becomes 2 [°C]		1384	1381	1384	1383	1383	1380

Table 6

		Example 17	Example 18	Example 19	Example 20	Example 21
glass composition [wt.%]	SiO ₂	68.3	67.9	70.0	69.0	69.5
	Al ₂ O ₃	2.87	2.85	1.70	2.40	1.90
	MgO	1.65	1.64	2.70	1.70	1.70
	CaO	9.23	9.16	8.04	9.11	9.00
	Na ₂ O	14.5	14.7	13.5	13.4	14.1
	K ₂ O	0.92	0.92	0.75	0.75	0.80
	Li ₂ O	0.50	0.33	1.50	2.00	1.00
	T-Fe ₂ O ₃	0.64	0.78	0.80	0.88	0.80
	FeO/T-Fe ₂ O ₃	0.18	0.27	0.26	0.24	0.24
	TiO ₂	0.16	0.29	0.06	0.10	0.06
optical properties	CeO ₂	1.17	1.44	0.65	0.60	0.87
	thickness[mm]	3.1	2.6	2.6	2.6	2.6
	YA[%]	77.2	75.0	75.7	74.9	76.5
	TG[%]	50.1	47.0	47.6	43.9	49.6
	Tuv[%]	15.3	12.7	15.6	11.8	15.0
	DW[nm]	498	503	499	515	508
physical properties	Pe[%]	3.0	2.6	2.6	2.7	2.4
	Tg [°C]	530.2	535.4	517.4	501.0	511.7
	Td [°C]	605.1	611.6	585.2	562.8	575.9
	$\alpha_{(50-350)} \times e^{-7}$	101.6	100.6	101.6	101.7	101.7
	E[GPa]	75.86	75.30	78.50	78.20	76.00
	surface compressive stress[MPa]	123.4	120.8	123.6	123.1	123.8
temperature at which log η becomes 2 [°C]		1380	1381	1378	1377	1378

[0071] It is apparent from Tables 5, 6 that each glass of the present invention has the temperature at which $\log \eta$ becomes 2, which is an index of high temperature viscosity, so that the ultraviolet/infrared absorbent green glass of the present 5 invention is excellent in meltability. Each ultraviolet/infrared absorbent green glass of the invention has also high surface compression, so that the ultraviolet/infrared absorbent green glass has large capacity for reinforcement.

[0072] Examples 11-21 have preferable optical properties, 10 such as: ultraviolet transmittance (Tuv) defined by ISO 9050 of not greater than 35%, dominant wavelength (DW) measured by using the CIE illuminant C of 490 to 560 nm and excitation purity (Pe) measured by using the CIE illuminant C of smaller than 6% when each glass has a thickness between 2.1 mm to 6 mm.

15 [0073] Examples 13-22 have more preferable optical properties to Examples 11 and 12 because ultraviolet transmittance (Tuv) of Examples 13-22 is lower than that of Examples 11 and 12.

[Comparative Examples 3 and 4]

[0074] The glass composition, the optical properties and the 20 physical properties of Comparative Examples 3 and 4 each of which was prepared in the same way as Examples 11-21 except for the glass composition are shown in Table 7.

Table 7

		Comparative Example 3	Comparative Example 4
glass composition [wt.%]	SiO ₂	71.3	69.9
	Al ₂ O ₃	1.55	1.48
	MgO	4.02	3.13
	CaO	7.99	8.49
	Na ₂ O	13.7	14.5
	K ₂ O	0.97	0.07
	Li ₂ O	0.00	0.00
	T-Fe ₂ O ₃	0.52	0.65
	FeO/T-Fe ₂ O ₃	0.22	0.31
	TiO ₂	—	0.14
optical properties	CeO ₂	—	1.65
	thickness[mm]	3.4	3.4
	YA[%]	82.9	75.4
	TG[%]	63.9	49.7
	Tuv[%]	32.2	10.0
	DW[nm]	500	510
physical properties	Pe[%]	1.7	2.2
	Tg [°C]	555.5	552.7
	Td [°C]	624.5	628.1
	$\alpha_{(50-350)} \times e^{-7}$	92.7	93.6
	E[GPa]	74.59	74.40
	surface compressive stress[MPa]	108.1	109.1
temperature at which log η becomes 2 [°C]		1410	1387

[0075] Comparative Examples 3 and 4 are conventional ultraviolet/infrared absorbent glasses which do not include Li₂O. These glasses have higher temperature at which log η becomes 2 and lower surface compression than those of the present 5 invention, resulting in reduction of productivity and capacity for reinforcement.

[0076] As described above, the present invention makes it possible to produce an ultraviolet/infrared absorbent glass having high visible light transmittance and greenish shade at 10 lower cost than that of conventional ones, because the present invention reduces viscosity of the glass at high temperature. Particularly, the ultraviolet/infrared absorbent green glass of the present invention exhibits superior capacity for reinforcement, so that it is especially fitted for a window 15 glass of a vehicle.